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Chapter 9 Review Stoichiometry Modern Chemistry Answers CHAPTER 9 REVIEW. Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_ The coefficients in a chemical equation represent the. (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a ...

## Chapter 9 Stoichiometry Review Answers

Modern Chemistry 77 Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. \_\_\_\_ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H

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## CHAPTER 9 REVIEW Stoichiometry

### CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS

Write the answer on the line to the left. Show all your work in the space provided.

- 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.
- 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the following equation:  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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CHAPTER 9 REVIEW stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.

Chapter 9 Review Stoichiometry Section 1 Answer Key

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