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Energy Levels □

Electrons in certain
energy levels □ Lower
energy levels: closer
to nucleus □ Higher
energy levels: further
from nucleus □

Ground State:
electron in lowest
energy level possible

Excited Atom □ Atom
has absorbed energy

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□ Excited state is unstable □ Atom soon emits same amount of energy absorbed □ energy seen as visible light Wave

Description of Light

Wavelength (λ):

distance between

corresponding points

...

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atom are described.

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wavelength,

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Bohr's Energy Levels.

1. electrons have

certain energy levels

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2. low energy levels-
closer to nucleus
3. high energy levels-
farther from nucleus
4. ground state-
electron in lowest
energy level possible.

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Energy Levels

Electrons in certain energy levels

- Lower energy levels: closer to nucleus
- Higher energy levels: further from nucleus

Ground State: electron in lowest energy level possible

Excited Atom: Atom has absorbed

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